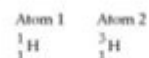


9) Which of the following best describes the relationship between the atoms described below?



- A) They are isomers.
- B) They are isotopes.
- C) They contain 1 and 3 protons, respectively.
- D) They each contain only 1 neutron.

Answer: B

Bloom's Taxonomy: Knowledge/Comprehension

Section: 2.2

10) The atomic number of nitrogen is 7. Nitrogen-15 has a greater mass number than nitrogen-14 because the atomic nucleus of nitrogen-15 contains \_\_\_\_\_.

- A) 7 neutrons
- B) 8 neutrons
- C) 8 protons
- D) 15 protons

Answer: B

Bloom's Taxonomy: Knowledge/Comprehension

Section: 2.2

11) From its atomic number of 15, it is possible to predict that the phosphorus atom has \_\_\_\_\_.

- A) 5 neutrons, 5 protons, and 5 electrons
- B) 15 neutrons and 15 protons
- C) 8 electrons in its outermost electron shell
- D) 15 protons and 15 electrons

Answer: D

Bloom's Taxonomy: Knowledge/Comprehension

Section: 2.2

12) Fluorine has an atomic number of 9. Which of the following would you do to a neutral fluorine atom to complete its valence shell?

- A) add 1 electron
- B) add 2 electrons
- C) remove 1 electron
- D) Nothing. If fluorine is neutral, it has a complete valence shell.

Answer: A

Bloom's Taxonomy: Application/Analysis

Section: 2.2

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