

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 10 REVIEW***States of Matter***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Identify whether the descriptions below describe an ideal gas or a real gas.
- \_\_\_\_\_ **ideal gas** a. The gas will not condense because the molecules do not attract each other.
- \_\_\_\_\_ **ideal gas** b. Collisions between molecules are perfectly elastic.
- \_\_\_\_\_ **real gas** c. Gas particles passing close to one another exert an attraction on each other.
2. The formula for kinetic energy is  $KE = \frac{1}{2}mv^2$ .
- a. As long as temperature is constant, what happens to the kinetic energy of the colliding particles during an elastic collision?  
**The energy is transferred between them.**
- b. If two gases have the same temperature and share the same energy but have different molecular masses, which molecules will have the greater speed?  
**Those with the lower molecule mass.**
3. Use the kinetic-molecular theory to explain each of the following phenomena:
- a. A strong-smelling gas released from a container in the middle of a room is soon detected in all areas of that room.  
**Gas molecules are in constant, rapid, random motion.**
- b. As a gas is heated, its rate of effusion through a small hole increases if all other factors remain constant.  
**As a gas is heated, each molecule's speed increases; therefore, the molecules pass through the small hole more frequently.**
4. a. \_\_\_\_\_ **b, d, c, a** List the following gases in order of rate of effusion, from lowest to highest. (Assume all gases are at the same temperature and pressure.)  
(a) He (b) Xe (c) HCl (d) Cl<sub>2</sub>

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