

Name: KEY

/ 15 Pts.

Chemistry Practice: Writing Chemical Formulas

Write a chemical formula for each substance.

- | | |
|--|---|
| 1. <u>NaCl</u> sodium chloride | 39. <u>$\text{Ca}(\text{NO}_3)_2$</u> calcium nitrate |
| 2. <u>N_2O_5</u> dinitrogen pentoxide | 40. <u>SO_3</u> sulfur trioxide |
| 3. <u>H_2S</u> hydrosulfuric acid | 41. <u>KCN</u> potassium cyanide |
| 4. <u>K_2SO_4</u> potassium sulfate | 42. <u>$\text{Pb}(\text{NO}_2)_2$</u> lead(II) nitrate |
| 5. <u>$\text{H}_2\text{C}_2\text{O}_4$</u> oxalic acid | 43. <u>H_2S</u> hydrogen sulfide |
| 6. <u>$\text{Ag}_2\text{C}_2\text{H}_3\text{O}_2$</u> silver acetate | 44. <u>CoCl_2</u> cobalt(II) chloride |
| 7. <u>$\text{Cr}(\text{NO}_3)_3$</u> chromium(III) nitrate | 45. <u>SF_6</u> sulfur hexafluoride |
| 8. <u>H_2CrO_4</u> chromous acid | 46. <u>Ca_3N_2</u> calcium nitride |
| 9. <u>$(\text{NH}_4)_2\text{CO}_3$</u> ammonium carbonate | 47. <u>CuI</u> copper(I) iodide |
| 10. <u>$\text{Ca}(\text{OH})_2$</u> calcium hydroxide | 48. <u>SiO_2</u> silicon dioxide |
| 11. <u>$\text{H}_2\text{C}_4\text{H}_4\text{O}_6$</u> tartaric acid | 49. <u>$\text{Sn}(\text{C}_2\text{H}_3\text{O}_2)_4$</u> tin(IV) acetate |
| 12. <u>$\text{Hg}_2(\text{NO}_3)_2$</u> mercury(II) nitrate | 50. <u>CCl_4</u> carbon tetrachloride |
| 13. <u>N_2O</u> dinitrogen monoxide | 51. <u>Cu_2S</u> cupric sulfide |
| 14. <u>Fe_2O_3</u> ferric oxide | 52. <u>$\text{Pb}_3(\text{PO}_4)_2$</u> lead(II) phosphate |
| 15. <u>$\text{Pb}(\text{ClO}_2)_2$</u> lead(II) chlorate | 53. <u>XeCl_4</u> xenon tetrachloride |
| 16. <u>$(\text{NH}_4)_3\text{PO}_4$</u> ammonium phosphate | 54. <u>Rb_2O</u> rubidium oxide |
| 17. <u>ZnCl_2</u> zinc chloride | 55. <u>MgSe</u> magnesium selenide |
| 18. <u>$\text{Ca}_3(\text{PO}_4)_2$</u> calcium phosphate | 56. <u>NH_4Cl</u> ammonium chloride |
| 19. <u>OF_2</u> oxygen difluoride | 57. <u>$\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_3$</u> iron(III) acetate |
| 20. <u>NaN_3</u> sodium azide | 58. <u>$\text{K}_2\text{Cr}_2\text{O}_7$</u> potassium dichromate |
| 21. <u>$\text{Fe}_2(\text{SO}_4)_3$</u> iron(III) sulfate | 59. <u>PBr_3</u> phosphorus tribromide |
| 22. <u>$\text{H}_2\text{As}_2\text{O}_5$</u> arsenous acid | 60. <u>Na_3PO_3</u> sodium phosphite |
| 23. <u>Cr_2O_3</u> chromium(III) oxide | 61. <u>Na_3PO_4</u> sodium phosphate |
| 24. <u>N_2O_4</u> dinitrogen tetroxide | 62. <u>$\text{Hg}_2(\text{NO}_3)_2$</u> mercury(II) nitrate |
| 25. <u>NH_4NO_3</u> ammonium nitrate | 63. <u>LiHCO_3</u> lithium bicarbonate |
| 26. <u>AuBr_3</u> gold(III) bromide | 64. <u>CrF_3</u> chromium(III) fluoride |
| 27. <u>CO</u> carbon monoxide | 65. <u>PbI_2</u> plumbous iodide |
| 28. <u>K_2CO_3</u> potassium carbonate | 66. <u>H_2SO_3</u> sulfurous acid |
| 29. <u>HIO_3</u> iodic acid | 67. <u>SnF_2</u> stannous fluoride |
| 30. <u>CsCl</u> cesium chloride | 68. <u>Hg_2CrO_4</u> mercuric chromate |
| 31. <u>$\text{Ni}(\text{MnO}_4)_2$</u> nickel(II) permanganate | 69. <u>KNO_3</u> potassium nitrate |
| 32. <u>$\text{Al}_2(\text{SO}_4)_3$</u> aluminum sulfate | 70. <u>SrCl_2</u> strontium chloride |
| 33. <u>$\text{Al}_2(\text{SO}_3)_3$</u> aluminum sulfite | 71. <u>P_4O_{10}</u> tetraphosphorus decoxide |
| 34. <u>$\text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$</u> barium acetate | 72. <u>KNO_3</u> potassium nitrate |
| 35. <u>$\text{Mn}(\text{OH})_3$</u> manganese(III) hydroxide | 73. <u>KNO_2</u> potassium nitrite |
| 36. <u>$\text{K}_2\text{H}_2\text{PO}_4$</u> potassium dihydrogen phosphate | 74. <u>K_3N</u> potassium nitride |
| 37. <u>HF</u> hydrofluoric acid | 75. <u>CaO</u> calcium oxide |
| 38. <u>$\text{Al}(\text{BrO}_3)_3$</u> aluminum bromate | 76. <u>$\text{Fe}(\text{IO}_4)_2$</u> iron(II) periodate |

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