

Name _____ Date _____ Class _____

CHAPTER 8 REVIEW*Chemical Equations and Reactions***SECTION 8-1****SHORT ANSWER** Answer the following questions in the space provided.

1. Match the symbol on the left with its appropriate description on the right.

- | | |
|----------------------------|--|
| _____ Δ | (a) A precipitate forms. |
| _____ \downarrow | (b) A gas forms. |
| _____ \uparrow | (c) A reversible reaction occurs. |
| _____ (l) | (d) Heat is applied to the reactants. |
| _____ (aq) | (e) A chemical is dissolved in water. |
| _____ \rightleftharpoons | (f) A chemical is in the liquid state. |

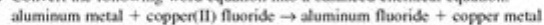
2. Finish balancing the following equation:



3. In each of the following formulas with coefficients, write the total number of atoms present.

- | | |
|-------|------------------------------------|
| _____ | a. 4SO_2 |
| _____ | b. 8O_2 |
| _____ | c. $3\text{Al}_2(\text{SO}_4)_3$ |
| _____ | d. $6 \times 10^{23} \text{HNO}_3$ |

4. Convert the following word equation into a balanced chemical equation:



5. One way to test the salinity of a water supply is to add a few drops of silver nitrate of known concentration to the water. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, leaving sodium nitrate in solution. Translate these sentences into a balanced chemical equation.

6. a. Balance the following equation: $\text{NaHCO}_3(s) \xrightarrow{\Delta} \text{Na}_2\text{CO}_3(s) + \text{H}_2\text{O}(g) + \text{CO}_2(g)$

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